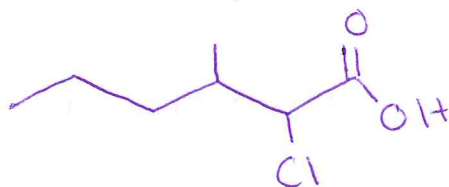


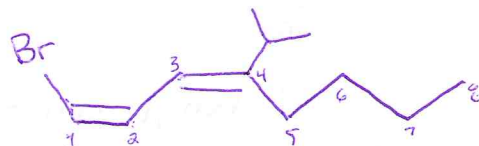
ORGANIC NAMING SHEET

1. Draw the following organic compounds. (Be sure you can draw all 3 types of diagrams)

a) 2-chloro-3-methylhexanoic acid



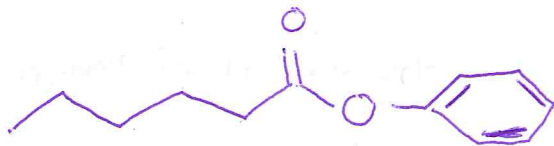
b) *cis*-1-bromo-4-isopropylocta-1,3-diene



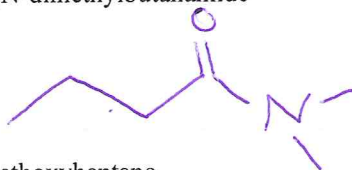
c) N-ethyl-N-methylpentan-3-amine



d) phenylhexanoate



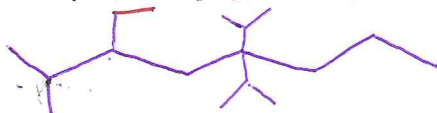
e) N,N-dimethylbutanamide



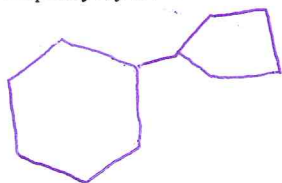
f) 3-ethoxyheptane



g) 3-ethyl-5,5-diisopropyl-2-methyloctane

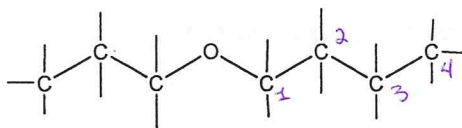


h) cyclopentylcyclohexane



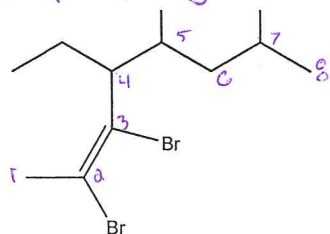
2. Name the following compounds.

a)

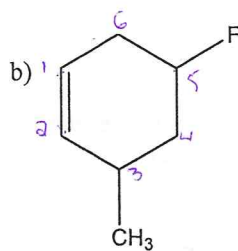


1-n-propoxybutane

d)

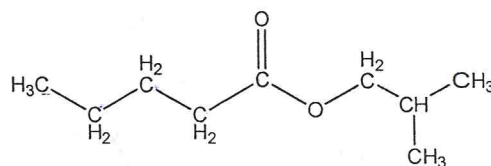


cis-2,3-dibromo-4-ethyl-5,7-dimethyloct-2-ene



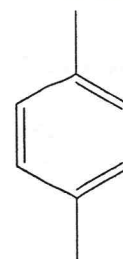
5-fluoro-3-methylcyclohex-1-ene

e)



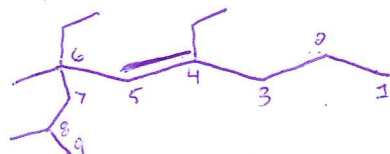
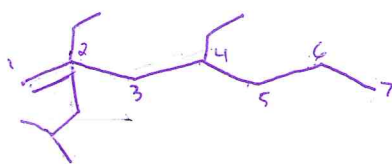
isobutyl pentanoate

1,4-dimethylbenzene
p-dimethylbenzene







3. Draw the following, correctly name it, and state in one sentence why the given name is incorrect, and there may be more than one reason: (yes, it is spelled as I intended, figure it out)

2-ethyl-2-isobutyl-4-ethyl-3-heptene



4,6-diethyl-6,8-dimethylnon-4-ene

4. For the following compounds, indicate the strength of the forces between the molecules. Some may need qualification like "stronger" or "weaker", use only if needed. Then, rate the order in which they would boil, starting from the lowest bp (1) to highest bp (4).

Compound	Intermolecular Forces	Ranking
 pentyne	<i>London-dispersion</i>	<i>1</i>
 Pentan-2-one	<i>London-dispersion, dipole-dipole</i> ^{<i>stronger</i>}	<i>3</i>
 Pentan-2-amine	<i>London-dispersion, dipole-dipole, H-bonding</i>	<i>4</i>
 1-methoxybutane	<i>London-dispersion, dipole-dipole</i>	<i>2</i>